**General Final Exam Review #3**

1. How many moles are there in 34 grams of copper (II) acetate? Copper (II) acetate is Cu(C2H3O2)2
2. How many grams are there in 56 moles of water?
3. If I have 77 moles of a gas at a temperature of 25o C and a pressure of 8.0 atm, what is the volume of this gas? R = 0.08206 Latm/molK.
4. If I were to further compress the gas from #3 to a volume of 10 liters, what would the pressure inside the container be?
5. Why do we assume that ideal gas molecules don’t have any volume?
6. What’s the molality of a solution that contains 30 grams of NaCl in 4.5 kilograms of water?
7. What’s the melting point of the solution from problem 6? (Kb = 0.512o C/m)